The phosphorescence of in aqueous solution is quenched by 1,1′-bis(2-sulfoethyl)-4,4′-bipyridinium inner salt (BSEP). Transient absorption attributable to BSEP⁻ ($\lambda_{max} \sim 610$ nm)[19] is observed in flash kinetic spectroscopic studies of aqueous solutions containing $Pt_2(P_2O_5)_4H_8^{4-}$ and BSEP, thereby establishing an electron-transfer quenching mechanism:

$$Pt_2(P_2O_5)_4H_8^{4-} + BSEP \xrightarrow{k_q} Pt_2(P_2O_5)_4H_8^{3-} + BSEP^-$$

Stern–Volmer analysis of the quenching yields $k_q = 5.5 \times 10^9$ M^{-1} s^{-1} ($[Pt_2(P_2O_5)_4H_8^{4-}] \sim 10^{-4}$ M; 0.1 M $NaClO_4$; 25 °C). Both the quenching reaction and the bimolecular back-electron-transfer ($k = 1 \times 10^9$ M^{-1} s^{-1} for $Pt_2(P_2O_5)_4H_8^{3-}$ and $BSEP^-$) are near the diffusion limit for such processes in aqueous solution at 25 °C.

The $^3A_{2u}(d\sigma^*p\sigma)$ state of $Pt_2(P_2O_5)_4H_8^{4-}$ is an extremely powerful one-electron reductant in aqueous solution. Preliminary experiments have shown that species such as $Os(NH_3)_5Cl^{2+}$ ($E_{1/2} = -1.09$ V vs. SCE)[20] and nicotinamide ($E_{1/2} = -1.44$ V vs. Ag/AgCl; CH_3OH, pH 7.2)[21] are readily reduced by $Pt_2(P_2O_5)_4H_8^{4-*}$. From these and related experiments it is apparent that $Pt_2(P_2O_5)_4H_8^{4-*}$ is a stronger reducing agent [$E°(3-/4-*) < -1$ V vs. NHE][22] than $Ru(bpy)_3^{2+*}$ ($E° = -0.88$ V vs. NHE)[23] in aqueous solution. We are now exploring several aspects of the photoredox chemistry of $Pt_2(P_2O_5)_4H_8^{4-}$, as the ease of generation of $Pt_2(P_2O_5)_4H_8^{4-*}$ suggests that it will be a useful reagent for a variety of substrate reductions.

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[19] Maverick, A. W. Ph.D. Thesis California Institute of Technology, 1982.

[20] Gulen, J.; Page, J. A. *J. Electroanal. Chem.* **1976**, *67*, 215–230.

[21] Meites, L.; Zuman, P.; Scott, W. J.; Campbell, B. H.; Kartos, A. M. "Electrochemical Data"; Wiley; New York, 1974; Part 1, AG85.

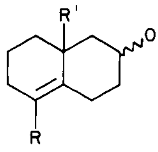
[22] It has not been possible to obtain a better estimate of $E°(4-*/3-)$ from spectroscopic and electrochemical measurements, because the electrochemical oxidation of $Pt_2(P_2O_5)_4H_8^{4-}$ in aqueous solution at 25 °C is not reversible (cyclic voltammetric measurements; graphite electrode; scan speed, 500 mV/s; $K_4[Pt_2(P_2O_5)_4H_8]$, 0.1 M $NaClO_4$).

[23] Navon, G.; Sutin, N. *Inorg. Chem.* **1974**, *13*, 2159–2164.

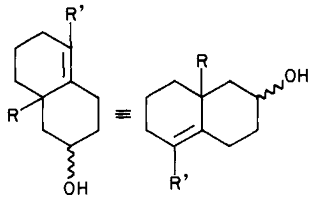
Additions and Corrections

On the Nonconcertedness of Allylic Cation Promoted π-Cyclization Reactions [*J. Am. Chem. Soc.* **1981**, *103*, 1285]. MLADEN LADIKA, IVO BREGOVEC, and DIONIS E. SUNKO,* Department of Chemistry, Faculty of National Sciences and Mathematics, University of Zagreb, 41000 Zagreb, Yugoslavia.

Page 1286: Structures **7** and **8** should be:



7



8

Stereochemical Analysis of γ-Replacement and γ-Elimination Processes Catalyzed by a Pyridoxal Phosphate Dependent Enzyme [*J. Am. Chem. Soc.* **1981**, *103*, 4921]. MICHAEL N. T. CHANG and CHRISTOPHER T. WALSH,* Departments of Chemistry and Biology, Massachusetts Institute of Technology, Cambridge, Massachusetts 02139.

Page 4925, Scheme VII: The last species is 4-*S*-2-ketobutyrate, not 4*R* as written.

Page 4925, 5th and 6th lines from the bottom, first column: The text should read "conversion of (*Z*)- and (*E*)-[4-^{2}H]-vinylglycines to (4*S*)- and (4*R*)-[^{1}H,^{2}H,^{3}H]-2-ketobutyrates, respectively, can be interpreted." The *S* and *R* designations were inadvertently transposed.

Page 4923, Table II, line 3: The numbers 1027.5 and 1029.3 are incorrectly transposed. The (*E*)-[4-^{2}H]-vinylglycine yields the 1029.3 sample of homoserine, the (*Z*)-vinylglycine yields the 1027.5 sample.

Synthesis, Structure, and Stability of (i,o)-Bicyclo[6.2.2]dodeca-9,11-dienes. Generation of Unusually Expanded Carbon–Carbon–Carbon Bond Angles [*J. Am. Chem. Soc.* **1981**, *103*, 215]. PAUL G. GASSMAN* and REBECCA C. HOYE, Department of Chemistry, University of Minnesota, Minneapolis, Minnesota 55455.

Page 217: In Table II for atom O2, the *z* value should be 0.4691 (1) instead of 0.4961 (1).

α-Disulfoxide and Sulfinic Anhydride in the Peroxy Acid Oxidation of 2-Methyl-2-propyl 2-Methyl-2-propanethiosulfinate [*J. Am. Chem. Soc.* **1981**, *103*, 6232–6235]. FILLMORE FREEMAN* and CHRISTOS N. ANGELETAKIS, Department of Chemistry, University of California, Irvine, California 92717.

Authors should read: Fillmore Freeman* and Christos N. Angeletakis, Department of Chemistry, University of California, Irvine, California 92717, and Tom J. Maricich, Department of Chemistry, California State University, Long Beach, California 90840.